

ACIDS, BASES AND SALTS

Full chapter notes | Board exam + school tests

FREE PREVIEW

- Acids and bases in daily life
- Indicator colour table
- One sample equation set
- List of what the full PDF contains
- Not the complete chapter — sample only

Written like a classroom teacher note

Acids, Bases and Salts

This chapter is scoring if you can write the **key equations**, read a **pH value**, and name the **use of each important compound** (NaOH, bleaching powder, baking soda, washing soda, plaster of Paris).

1. Acids and bases around us

Acids taste sour. Bases taste bitter and feel soapy. In Class 10 we study their chemical behaviour, not only taste (tasting lab chemicals is never allowed).

- Natural acids: citric acid (lemon), acetic acid (vinegar), lactic acid (curd), tartaric acid (tamarind), oxalic acid (tomato).
- Mineral acids in the lab: HCl, H_2SO_4 , HNO_3 .
- Bases: NaOH (soap making), KOH, $\text{Mg}(\text{OH})_2$ (milk of magnesia), $\text{Ca}(\text{OH})_2$ (whitewash).

Teacher note: Acidic / basic nature is due to ions in water. Dry HCl gas does not turn blue litmus red. An aqueous solution does, because H^+ / H_3O^+ ions form.

2. Indicators

An indicator is a dye that shows different colours in acid and base.

Indicator	In acid	In base
Blue litmus	Red	Stays blue
Red litmus	Stays red	Blue
Methyl orange	Red	Yellow
Phenolphthalein	Colourless	Pink
Turmeric	Stays yellow	Reddish brown
Red cabbage / china rose	Pink-red / dark pink	Green / green

Olfactory indicators change smell in acid or base (onion, vanilla, clove oil). Used less in papers than litmus / phenolphthalein.

3. Sample equations (full set is in the paid PDF)



Locked in the full version

- All acid–base reactions with tests for H_2 and CO_2
- pH scale and tooth decay / soil / acid rain
- Chlor-alkali, bleaching powder, baking soda, washing soda, PoP
- Water of crystallisation
- 16 practice questions + answer key

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